

Use the Periodic Table to complete questions #2 and #3.

2a. How can you determine the number of protons, neutrons, and electrons from the Periodic Table?

Protons: _____

Neutrons: _____

Electrons: _____

2b. Determine the number of protons, neutrons, and electrons for your element by using the atomic number and atomic mass.

What is the element's atomic number? _____

What is the element's atomic mass? _____

Protons = _____ Neutrons = _____ Electrons = _____

(NOTE: Atomic mass - atomic number = neutrons)

3. Determine the number of electrons in each electron energy level.
(Level 1 = 2, Level 2 = 8, Level 3 = 18, Level 4 = 32)

Number of electrons in your element:

Level 1 = _____ Level 2 = _____ Level 3 = _____ Level 4 = _____

Conclusion:

1a. In what ways is your model of an atom different from an actual atom? _____

1b. How is your model of an atom like an actual atom? _____

2. Why do the different elements have different numbers of energy levels? _____

3. If an atom gained a neutron in its nucleus (became an isotope), predict —

a) Would the atomic number change? _____

b) Would the atomic mass change? _____

c) Would it become (a totally different element, or a different form of the same element)

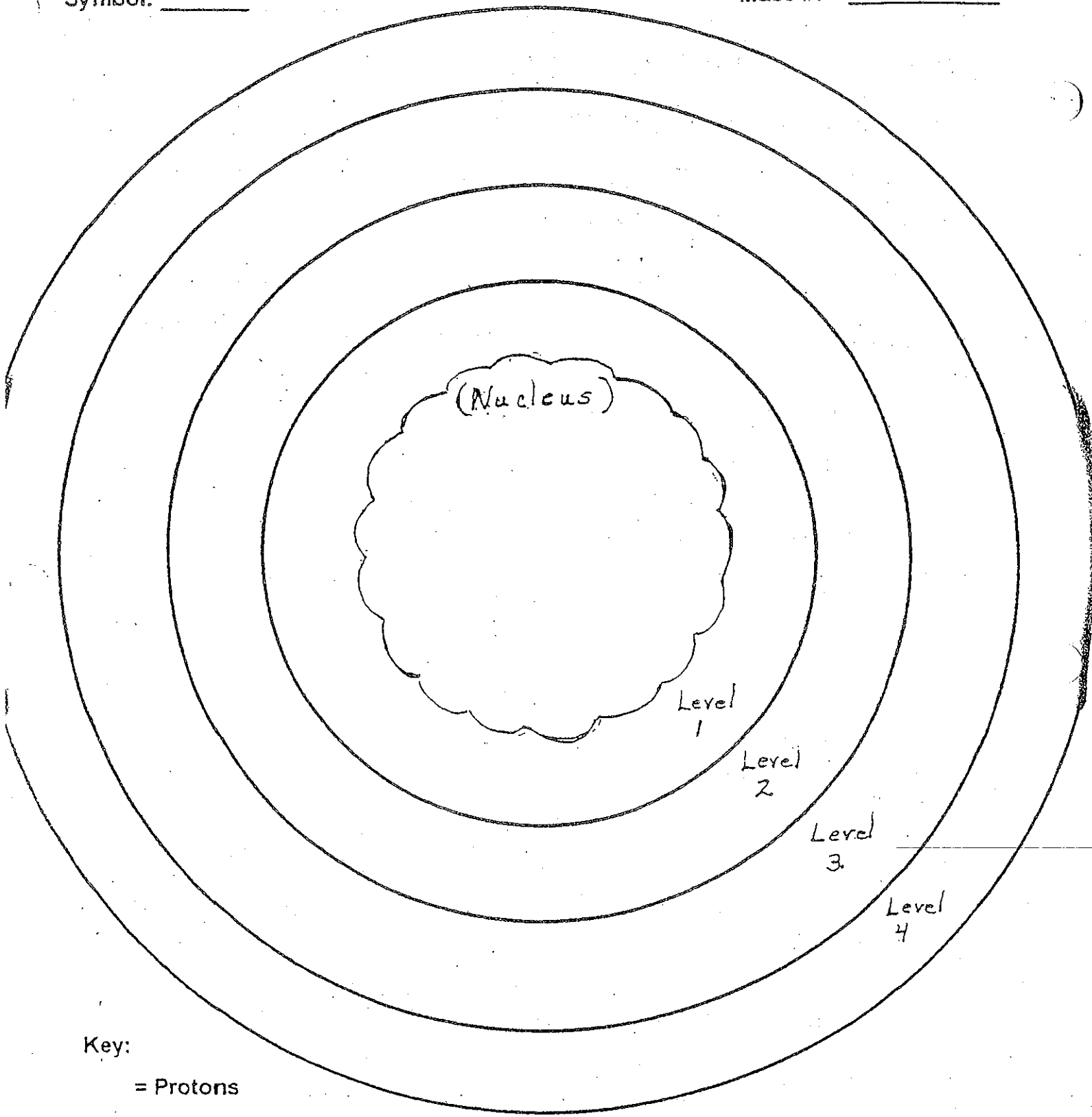
Circle one

Element: _____

Atomic #: _____

Symbol: _____

Mass #: _____



Key:

= Protons

= Neutrons

= Electrons

Name _____